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3. a chemical
compound always
contains the.... mass
cannot be created or
destroyed in.... when
two elements combine
to form two o....
subatomic particle with
a negative char.... law
of definite proportion.
law of conservation of
mass. law of multiple
proportions. electron.

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CHAPTER 3 REVIEW

Atoms: The Building
Blocks of Matter

SECTION 3 SHORT

ANSWER Answer the
following questions in
the space provided. 1.
Explain the difference
between the mass
number and the atomic
number of a nuclide.

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Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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Test. Honors Chemistry
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PLAY. law of
conservation of mass.
mass is neither created
nor destroyed during
ordinary chemical
reactions or physical
changes $2H_2 + O_2 = 2H_2O$
O. law of definite
proportions.

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State the following measured quantities in the units indicated.

- a. 5.2 cm of magnesium ribbon in millimeters
- b. 0.049 kg of sulfur in grams
- c. 1.60 mL of ethanol in microliters
- d. 0.0025 g of vitamin A in micrograms
- e.

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0.020 kg of tin in
milligrams f. 3 kL of
saline solution in liters

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Name CHAPTER 3 TEST
continued Date Class
FILL IN THE BLANK

Write the correct term
(or terms) in the space
provided. 9, If a
particular compound is
composed of elements
A and B, the ratio of

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Answers
the mass of B to the mass of A will always be the same. This is a statement of the law of conservation of mass. exactly 12 g of carbon-12 is referred to as a(n) 11.

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Test Chapter Test B,
continued 16. The
measure of the ability
of an atom in a
chemical compound to
attract electrons from
another atom in the
compound is called
_____.

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Test Answer Key

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