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Chapter 14 Properties Of Gases

states that, at constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressure of the component gases

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Properties Of

diffusion is the tendency of molecules to move toward areas of lower concentration until the concentration is uniform throughout

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Chapter 14 The Behavior of Gases 147
SECTION 14.1
PROPERTIES OF GASES (pages 413–417)
This section uses kinetic theory to

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explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature.

Compressibility (pages 413-414) 1. Look at Figure 14.1 on page 413.

SECTION 14.1 PROPERTIES OF GASES(pages 413-417)

14.1 Properties of

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Properties Of

Gases In organized soccer, there are rules about equipment. For international competitions, the ball's mass must be not more than 450 grams and not less than 410 grams. The pressure of the air inside the ball must be no lower than 0.6 atmospheres and no higher than 1.1 atmospheres at sea level. A ball that is

14.1 Properties of

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420 Chapter 14 Gases

Almost all the volume of a gas is empty space. Gases can be compressed by moving gas particles closer together because of this low density of particles. • Gas particles are in constant, random motion. Gas particles spread out and mix with each other because of this motion. The particles move in

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Chapter 14: Gases

Chapter 14 - The
Behavior of Gases -

14.1 Properties of
Gases - 14.1 Lesson

Check - Page 454: 6

Answer If the

temperature is

constant, quadrupling

the volume would

cause the pressure of

an enclosed gas to be

reduced to one quarter

of its original value.

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Chapter 14 - The Behavior of Gases - 14.1 Properties of ...

SECTION 14.1

PROPERTIES OF GASES

(pages 413–417) This section uses kinetic theory to explain the properties of gases.

This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature.

SECTION 14.1

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413-417) -

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Chapter 14 (& 13.1)
- Gases - KEIO
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Properties of Gases.
Prentice-Hall Chapter
14.1. Dr. Yager.
Objectives Explain why
gases are easier to

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compress than solids
or liquids Describe the
three factors that
affect gas pressure

Compressibility

Compressibility is a
measure of how much
the volume of matter
can decrease under
pressure. Examples of
Compressibility Car air
bags A person during a
collision compresses
the gas inside the bag
to absorb energy.

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432 Chapter 14 Table

14.1 Composition of Dry Air Component	Volume (%)	Partial pressure (kPa)
Nitrogen	78.08	79.11
Oxygen	20.95	21.22
Carbon dioxide	0.04	0.04
Argon and others	0.93	0.95
Total	100.00	101.32

14.4 Gases:

Mixtures and

Movements The top of
Mount Everest is more
than 29,000 feet above
sea level. A list of gear

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Properties Of Gases Answers

for an expedition to

14.4 Gases: Mixtures and Movements

Chemistry: Gases

Chapter 14. Define

Boyle's Law. Boyle's

Law Equation. Define

Charles's Law.

Charles's Law

Equation. At constant

temperature and mass

(moles), the pressure is

inve.... $P_1V_1 = P_2V_2$. At

constant pressure and

mass (moles), the

volume of a gas is....

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Properties Of
 $V_1/T_1 = V_2/T_2.$

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easily compressed. The volume of the particles in a gas, such as. air, is small compared to the overall volume of. the gas. So, the distance between particles in.

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air is much greater than the distance between particles in a solid, such as wood. Under pressure, the particles in a gas can be forced.

Chapter 14

With effusion and diffusion, the type of particle is important: Gases of lower molar mass diffuse and effuse faster than gases of higher molar mass. Helium effuses

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and diffuses faster
than nitrogen.

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14. In an inverse relationship, the ratio of two variable quantities is constant.

15. When using the combined gas law, pressure must always be in kilopascals but temperature can be in kelvins or degrees Celsius. 273

16. When

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20.0 i of O_2 is warmed from -30.00°C to 85.00°C at constant pressure, the new volume is 29.5 L.

Vocabulary Boyle's law

Bozeman Public Schools

End-of-Chapter
Material; Chapter 6.
Gases. Introduction to
Gases; Pressure; Gas
Laws; Other Gas Laws;
The Ideal Gas Law and
Some Applications; Gas
Mixtures; Kinetic

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Properties Of
Gases; Molecular
Theory of
Gases; Molecular
Effusion and Diffusion;
Real Gases; End-of-
Chapter Material;
Chapter 7. Energy and
Chemistry. Introduction
to Energy and
Chemistry; Formation
...

The Ideal Gas Law and Some Applications - Introductory ...

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Behavior of Gases 147
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Properties Of Gases Answers SECTION 141 PROPERTIES OF GASES(pages 413-417)

This section uses kinetic theory to explain the properties of gases This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature

Compressibility (pages 413-414) 1 Look at Figure 141 on page 413

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Chapter 5 Properties of
Gases Chemistry

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Chapter 5 Properties of Gases

In this lesson, we will
discuss the many

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characteristics of gases and how knowing the microscopic properties of gas particles will help you understand the macroscopic properties of a gas.

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